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(21) International Application Number: PCT/IE98/00008 (22) International Filing Date: 6 February 1998 (06.02.98) (30) Priority Data: S970082 6 February 1997 (06.02.97) IE (71) Applicant (for all designated States except US): UNIVERSITY COLLEGE DUBLIN [IE/IE]; Belfield, Dublin 4 (IE). (72) Inventors; and (75) Inventors/Applicants (for US only): FITZMAURICE, Donald [IE/IE]; 10 Lansdowne Court, Haddington Road, Ballsbridge, Dublin 4 (IE). NAGARAJA, Rao, S. [IE/IE]; 3 St. Gaten's Court, Whitechurch Road, Rathfarnham, Dublin 14 (IE). CINNSEALACH, Rachel [IE/IE]; 198 Garran Cluana, Dun-Droma, Baile Atha Cliath 16 (IE). ENRIGHT, Brendan [IE/IE]; 48 Pinewood, Elm Park, Wilton, Cork (IE). (74) Agent: DUFFY, Assumpta; F.R. Kelly & Co., 27 Clyde Road, Ballsbridge, Dublin 4 (IE).		(81) Designated States: AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, GM, GW, HU, ID, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG). Published <i>With international search report.</i> <i>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>
(54) Title: ELECTROCHROMIC SYSTEM		
(57) Abstract The present invention relates to an electrochromic system comprising a first electrode disposed on a transparent or translucent substrate and a second electrode, an electrolyte, an electron donor and a nanoporous-nanocrystalline film of a semiconducting metallic oxide having a redox chromophore adsorbed thereto being intermediate the first and second electrodes. Such a system has a very rapid electrochromic effect.		

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- 1 -

ELECTROCHROMIC SYSTEM

The present invention relates to improvements in or relating to electrochromic systems.

5

As the energy performance of buildings and automobiles becomes an increasingly important design feature, strategies for optimising performance in this respect are receiving considerable attention.

10

An important aspect of the energy performance of the above relates to the incident radiation transmitted by the window area of a building. Such concerns are further complicated by the need to ensure occupant comfort. It is in this context that the electrochromic (EC) window technology has assumed increasing importance, the amount of incident radiation transmitted by such windows being electronically controllable. Effective implementation of EC window technology in buildings is expected to provide the following benefits:

1. Reduce adverse cooling effects.
 - Reduce cooling energy.
- 25 2. Down-size air conditioning plant.
 - Reduce peak electricity demand.
3. Increase beneficial effects of daylight.
 - Reduce lighting energy.
- 30 4. Reduce peak electricity demand.
5. Increase occupant comfort.
 - Increase thermal comfort.
 - Increase visual comfort.

- 2 -

Even greater benefits would be expected to accrue in an automobile, where the ratio of glazed surface to enclosed volume is significantly larger than in a typical building. Specifically, effective
5 implementation of EC window technology in automobiles is expected to provide the following benefits in addition to those in the built environment :

1. Increased motoring safety.
10 Reduced glare.
 Mirror control.
 Head-up display.

EC technology is not limited to the applications
15 described above. Others include privacy glass, angle-independent high-contrast large-area displays, glare-guards in electronic devices, electronic scratch pads.

20 Existing EC devices, including those commercially available, are non-optimal for the large glazing areas encountered in building and automotive applications and are based on technologies which are process and energy intensive. Therefore new EC technologies, resulting in
25 improved device specification and which may be manufactured more easily at a lower cost, will be commercially important. It is noted in this context, that the current market for EC window technologies in buildings and automobiles is estimated world-wide at
30 over \$2 billion.

For an overview of these and related topics see the review *Large-Area Chromogenics : Materials and Devices for Transmittance Control* (Eds. Lampert and
35 Grangvist), SPIE Institutes for Advanced Optical

- 3 -

Technologies Series Vol. 4. Existing EC devices are found in one of the two categories outlined below. Firstly, there are those devices based on ion insertion reactions at metal oxide electrodes. To ensure the
5 desired change in transmittance the required number of ions must be intercalated in the bulk electrode to compensate the accumulated charge. However, use of optically flat metal oxide layers requires bulk
10 intercalation of ions as the surface area in contact with electrolyte is not significantly larger than the geometric area. As a consequence the switching times of such a device are typically of the order of tens of seconds.

15 Secondly, there are those devices based on a transparent conducting substrate coated with a polymer to which is bound a redox chromophore. On applying a sufficiently negative potential there is a transmittance change due to formation of the reduced
20 form of the redox chromophore. To ensure the desired change in transmittance a sufficiently thick polymer layer is required, the latter implying the absence of an intimate contact between the transparent conducting substrate and a significant fraction of the redox
25 chromophores in the polymer film. As a consequence the switching times of such a device are, as above, typically of the order of tens of seconds.

30 It is an object of the present invention to provide an improved EC system.

According to the invention there is provided a nanoporous-nanocrystalline film comprising a semiconducting metallic oxide having a redox
35 chromophore adsorbed thereto.

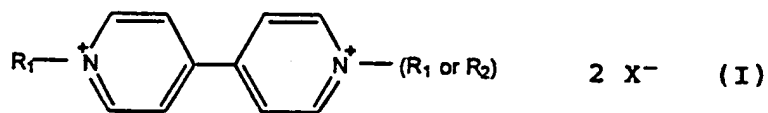
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- A "nanocrystalline film" is constituted from fused nanometer-scale crystallites. In a "nanoporous-nanocrystalline" film the morphology of the fused nanocrystallites is such that it is porous on the nanometer-scale. Such films, which may hereinafter be referred to as nanostructured films, typically possess a surface roughness of about 1000 assuming a thickness of about 10 μm .
- 10 The nanostructured films used in the present invention colour on application of a potential sufficiently negative to accumulate electrons in the available trap and conduction band states. As a consequence of the high surface roughness of these films, ions are readily
- 15 adsorbed/intercalated at the oxide surface permitting efficient charge compensation and rapid switching, i.e. the need for bulk intercalation is eliminated. However, despite the rapid switching times in such films, the associated change in transmittance is not sufficient
- 20 for a commercial device. To overcome this limitation a redox chromophore is adsorbed at the surface of the transparent nanostructured film which, when reduced, increases the extinction coefficient of an accumulated trapped or conduction band electron by more than an
- 25 order of magnitude. Further, due to the nanoporous structure and associated surface roughness of the nanocrystalline films used, the redox chromophore is effectively stacked as in a polymer film, while at the same time maintaining the intimate contact with the
- 30 metal oxide substrate necessary to ensure rapid switching times.

The redox chromophore may be any suitable redox chromophore and preferably comprises a compound of the

35 general formula I

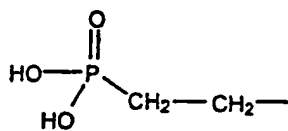
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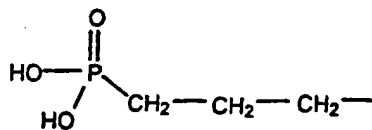
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wherein X is a charge balancing ion such as a halide; R₁ is any one of the following:

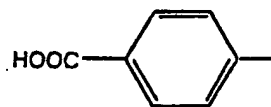
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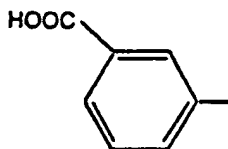
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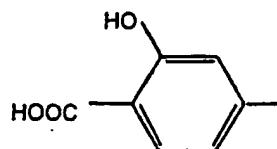
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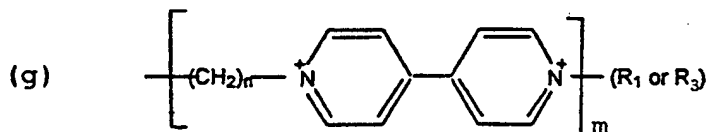
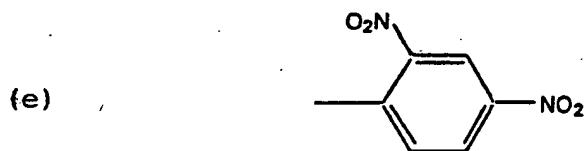
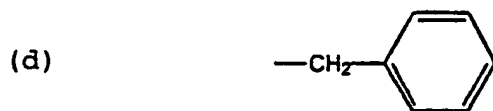


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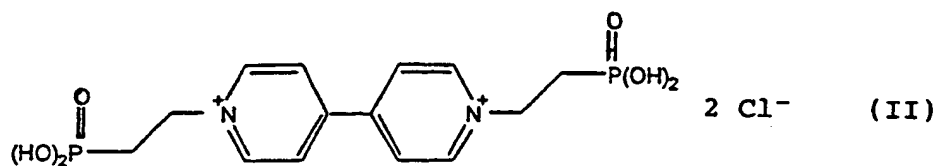
R₂ is any one of the following:

- 6 -



wherein R_1 is as defined above, R_3 is any of the formulae (a) to (f) given above under R_2 , m is an integer of from 1 to 6, preferably 1 or 2 and n is an integer of from 1 to 10, conveniently 1 to 5.

A particularly preferred redox chromophore is a compound of formula II, viz.
bis-(2-phosphonoethyl)-4,4'-bipyridinium dichloride



- 7 -

The semiconducting metallic oxide may be an oxide of any suitable metal, such as, for example, titanium, zirconium, hafnium, chromium, molybdenum, tungsten, vanadium, niobium, tantalum, silver, zinc, strontium, iron (Fe^{2+} or Fe^{3+}) or nickel or a perovskite thereof. TiO_2 , WO_3 , MoO_3 , ZnO and SnO_2 are particularly preferred.

The invention also provides an electrochromic system comprising: a first electrode disposed on a transparent or translucent substrate; a second electrode; an electrolyte; an electron donor; and an electrochromic layer comprising a nanoporous-nanocrystalline film according to the invention intermediate the first and second electrodes.

The substrate is suitably formed from a glass or a plastics material. Glass coated with a conducting layer of fluorine doped tin oxide or indium tin oxide is conveniently used in the EC system of the present invention.

The electrolyte is preferably in liquid form and preferably comprises at least one electrochemically inert salt optionally in molten form in solution in a solvent. Examples of suitable salts include hexafluorophosphate, bis-trifluoromethanesulfonate, bis-trifluoromethylsulfonylamidure, tetraalkylammonium, dialkyl-1,3-imidazolium and lithium perchlorate. Examples of suitable molten salts include trifluoromethanesulfonate, 1-ethyl,3-methyl imidazolium bis-trifluoromethylsulfonylamidure and 1-propyl-dimethyl imidazolium bis-trifluoromethylsulfonylamidure. Lithium perchlorate is particularly preferred.

- 8 -

The solvent may be any suitable solvent and is preferably selected from acetonitrile, butyronitrile, glutaronitrile, dimethylsulfoxide, dimethylformamide, dimethylacetamide, N-methyloxazolidinone,
5 dimethyl-tetrahydropyrimidinone, γ -butyrolactone and mixtures thereof.

The electron donor is preferably a metallocene or a derivative thereof. The electron donor is preferably
10 soluble in the electrolyte solvent. Ferrocene is particularly preferred.

The invention is illustrated in the following Example.

15

EXAMPLE

(a) A 2.5 cm x 2.5 cm transparent nanostructured film, consisting of a 4 μ m thick layer of fused TiO₂ nanocrystallites, was deposited on a 3.3 cm x 3.3 cm
20 fluorine doped tin oxide on glass substrate (Glastron, Trade Mark). A colloidal TiO₂ dispersion was prepared by hydrolysis of titanium tetraisopropoxide. The average diameter of the initially formed crystallites (7 nm) was increased by autoclaving at 200°C for 12
25 hours to 12 nm. Concentrating the autoclaved dispersion to 160 g/l and adding Carbowax (Trade Mark) 20000 (40% wt. equiv. of TiO₂) yielded a white viscous sol. (Carbowax 20000 is an ethylene glycol polymer whose average molecular weight is 20000.) A 4 μ m thick
30 layer of the above sol was deposited using a screen printing technique on the conducting glass substrate. The resulting gel-film was dried in air for 1 h, sintered at 450°C for 12 h and stored in a darkened vacuum desiccator prior to use.

- 9 -

(b) A redox chromophore, bis-(2-phosphonoethyl)-4,4'-bipyridinium dichloride was prepared by adding 4,4'-bipyridine (4.4 g) and diethyl-2-ethylbromophosphonate (15.0 g) to water (75 ml). The reaction mixture was refluxed for 72 h and allowed to cool. Following addition of conc. hydrochloric acid (75 ml) the reaction mixture was refluxed for a further 24 h. To recover the product, the reaction mixture was concentrated to 50 ml, isopropyl alcohol (200 ml) added drop-wise, stirred on ice for one hour and filtered. The white crystalline product was washed with cold isopropyl alcohol and air dried to give pure bis-(2-phosphonoethyl)-4,4'-bipyridinium dichloride (12.72 g, 84.24 % yield). Calculated for bis-(2-phosphonoethyl)-4,4'-bipyridinium dichloride ($C_{14}H_{20}N_2Cl_2O_6P_2$) : C, 37.77; H, 4.53; N, 6.29. Found: C, 35.09; H, 4.49; N, 6.09. 1H NMR (water- d_2): δ 2.31-2.43 (m, 4H); δ 4.68-4.80 (m, 4H); δ 8.33 (d, unresolved metacoupling, 4H); δ 8.94 (d, unresolved metacoupling, 4H).

(c) TiO_2 films, prepared as described above, were modified by adsorption of the redox chromophore prepared above from an aqueous solution (0.02 mol.dm^{-3}) over 24 h, washed with distilled deionised water, dried in air and stored in a darkened vacuum desiccator for 48 h prior to use.

(d) Using a screen printing technique, a 0.25 cm border of a proprietary epoxy resin (Araldite, Trade Mark) was deposited on a second 3.3 x 3.3 cm fluorine doped tin oxide conducting glass, leaving a small opening in one corner. This piece of conducting glass was placed on top of the modified TiO_2 film prepared as described above and left to set for 24 h.

- 10 -

(e) To complete construction of the EC system, the above sandwich structure was back-filled using an argon pressure with an electrolyte solution consisting of LiClO_4 ($0.05 \text{ mol} \cdot \text{dm}^{-3}$) and ferrocene ($0.05 \text{ mol} \cdot \text{dm}^{-3}$) in γ -butyrolactone (m.p. -45°C , b.p. 204°C). The components of the electrolyte solution were carefully purified and rigorously dried prior to use. The opening was subsequently closed using Araldite (Trade Mark).

10

The EC system prepared as described in the Example is illustrated in Figures 1-3 of the accompanying drawings.

15 Figure 1 is a schematic view of the prepared film disposed on a substrate;

Figure 2 is a schematic view of the prepared electrochromic system including the film shown in
20 Figure 1; and

Figure 3 is an exploded view of the electrochromic system of Figure 2.

25 With reference to the drawings and in particular to Fig. 1, there is shown a first glass substrate 11 having a conductive layer 13 of fluorine doped tin oxide coated thereon. The exposed surface of the layer 13 is coated with a transparent nanostructured film 14
30 of TiO_2 having a redox chromophore 15 adsorbed thereon. The redox chromophore 15 is bis-(2-phosphonoethyl)-4,4'-bipyridinium dichloride prepared as described in the Example. Figs. 2 and 3 illustrate an EC system 10 according to the invention comprising the first glass
35 substrate 11 with the layer 13 and the modified TiO_2

- 11 -

film 14 shown in Fig. 1 and a second glass substrate 22 having a conductive layer 23 of fluorine doped tin oxide coated thereon. The second glass substrate 22 has a 0.25 cm border 24 of epoxy resin deposited thereon with a small gap 25, which is sealed after addition of the electrolyte/electron donor solution 16 described above.

It will be observed that construction of the EC system 10 according to the invention is simple and utilises low-cost and non-toxic materials. These are particularly attractive features in the context of the large-scale manufacture of the EC system 10.

It should also be noted that due to surface roughness, of the order of 500 for a 4 μm film, no spacer is required in an EC system of the invention.

In prior art electrochromic systems, a dielectric spacer must be included to isolate the electrodes electrically from each other. In the present invention, no such spacer is required because the solid particle nature of the nanocrystalline film provides for sufficient electrical isolation between the electrodes. In a commercial version of the EC system according to the invention, the absence of a spacer will have a positive impact on the manufacturing costs of the system.

A number of EC systems prepared as described in the above Example were tested by applying 10 000 cycles (15s at -1.00V and 15s at +1.0V at room temp.). A typical set of test results is shown in Figs. 4a and 4b.

- 12 -

Specifically, shown in Fig. 4a are the absorption spectra in the low transmittance (LT) state, after 1 and 10 000 cycles. It will be observed that this spectrum, as expected, corresponds to that of the radical cation of the viologen moiety of the redox chromophore. It will also be noted that, in practice, this corresponds to an intense blue coloration of the EC system and that the extent of this coloration is not diminished after 10 000 cycles.

Concerning the change in transmittance, this is conventionally represented as in Fig. 4b. Specifically, it will be observed that the transmittance decreases from about 70% (point a, 10 000 cycles) to about 8% (point b, 10 000 cycles). Significantly, the transmittance has decreased to less than 20% of its initial value in less than 1 s. Similar behaviour is observed upon switching the EC system to the high transmittance (HT) state.

As stated, there is no significant degradation in performance over 10 000 cycles. This is quantitatively demonstrated in Figs. 5a and 5b. Specifically, shown in Fig. 5a is the change in transmittance measured after 1, 10, 100, 1000 and 10 000 cycles. It should be noted that the slight improvement in performance observed is a reproducible feature. Shown in Fig. 5b, are plots of the switching times (as defined above) for the same EC system. These are consistently between 0.9 s and 1.1 s.

Nanostructured TiO_2 films were deposited on the following conducting glass substrates : Indium tin oxide glass and fluorine doped tin oxide glass. No significant difference in the performance in the

- 13 -

resulting EC system was detected. The time for which a film is fired is important for the following reason:

If a film is fired for 1 h its porosity, and consequently its surface roughness, will be optimal.

5 However, under the same conditions, film conductivity will be less than optimal due to incomplete sintering of the constituent nanocrystallites. Conversely, if a film is fired for 168 h, its connectivity, and consequently its conductivity, will be optimal.

10

However, under the same conditions, film porosity will be less than optimal due to collapse of the film's nanostructure as will be observed in Figs. 6a and 6b.

Shown in Fig. 6a are the transmittance changes after
15 10 000 cycles on switching an EC system in which the constituent nanoporous-nanocrystalline film has been fired for the indicated time. The best performance is obtained for systems containing films that have been fired for 12 h. However, as can be seen from Fig. 6b,
20 while there is improved colouring on going from 6 to 12 h firing time there is no corresponding decrease in the colouring or clearing time. Film thickness was 4 μm or less.

25 The film firing temperature should be above about 400°C to remove the added Carbowax, the addition of which is essential to ensure a porous film, and less than 500°C to prevent conversion of anatase to rutile, the latter being a significantly poorer conductor. For these
30 reasons the firing temperature was fixed at approximately 450°C.

The substituent groups of the redox chromophore are irreversibly chemisorbed at Ti^{4+} sites at the surface
35 of the TiO_2 nanocrystallites that constitute the

- 14 -

nanoporous-nanocrystalline film. These substituent groups, referred to as linker groups, serve, therefore, to irreversibly attach the redox chromophore to the surface of the nanoporous-nanocrystalline film. The density of these states (about $5 \times 10^{13} \text{ cm}^{-2}$), and the surface roughness, (about 1000 for a $4 \mu\text{m}$ film) provide the upper limit for the number of molecular amplifiers which may be adsorbed per unit geometric area. It should also be noted that, unlike previous linkers, there is no discoloration of the modified film due to the existence of a charge transfer interaction between the occupied molecular orbitals of the linker and the available conduction band states of the semiconductor substrate. With regard to the redox chromophore, the viologen moiety is stable with a large associated change in extinction for a one electron reduction. Further, the redox chromophore may be readily modified to change its electrochemical and optical properties by use of the various substituents associated with R in the general formula. Each variation possesses different formal potentials and different colours upon being switched. Furthermore, the redox chromophore may be readily prepared with high yield in a pure form and, perhaps most importantly, adsorbed onto the TiO_2 substrate from an aqueous solution.

One parameter which was studied in respect of the redox chromophore was the extent of modifier adsorption in a given period. As would be expected, the redox chromophore is adsorbed to an increasing extent from more concentrated solutions in a shorter time. In practice, for a 0.02 mol dm^{-3} aqueous solution of the redox chromophore, close to maximum coverage is observed after about 6 h with only a small subsequent increase in coverage during the following week, see

- 15 -

Figs. 7a and 7b. Some variability of this process is observed. The electrolyte solution consists of LiClO_4 (0.05 mol.dm^{-3}) and ferrocene (0.05 mol.dm^{-3}) in γ -butyrolactone (BL) (m.p. -45°C , b.p 204°C). The
5 concentration of the LiClO_4 and ferrocene were systematically varied and the results of these studies are summarised in Figs. 8a, 8b, 9a and 9b.

The concentration of added LiClO_4 , in the range
10 0.05 mol.dm^{-3} to 0.20 mol.dm^{-3} , has no effect on the magnitude of the transmittance change or on the colouring or clearing times (see Figs. 8a and 8b). On the other hand increasing the concentration of added ferrocene, in the range 0.05 mol.dm^{-3} to 0.20 mol.dm^{-3} ,
15 increases significantly the magnitude of the transmittance change and less significantly the colouring and clearing times (see Figs. 9a and 9b). The disadvantage of the latter is that the ferrocene attacks the epoxy resin seal on the cell and results in
20 device failure after about 48 h.

- 16 -

The legends to Figures 4-9 are as follows:

Figure 4 : (a) Absorption spectrum of the EC system 10
in low transmittance state. (b) Test result of
5 modified EC system 10 in (a) after 1 and 10 000 test
cycles.

Figure 5 : (a) Change in transmittance at 600 nm of the
EC system 10 in Fig.4 during 10 000 test cycles. (b)
10 Change in colouring and clearing times of the EC system
10 in (a) during 10 000 test cycles.

Figure 6 : (a) Change in transmittance at 600 nm of
modified EC system 10 after 10 000 test cycles as a
15 function of the firing time of nanostructured film. (b)
Test results of modified EC system 10 in (a) after
10 000 test cycles.

Figure 7 : (a) Change in transmittance at 600 nm of
20 modified EC system 10 after 10 000 test cycles as a
function of the dying time of nanostructured film. (b)
Test results of modified EC system 10 in (a) after
10 000 test cycles.

25 Figure 8 : (a) Change in transmittance at 600 nm of
modified EC system 10 containing $0.20 \text{ mol} \cdot \text{dm}^{-3} \text{ LiClO}_4$
during 10 000 test cycles. (b) Change in colouring and
clearing times of modified EC system 10 in (a) during
10 000 test cycles.

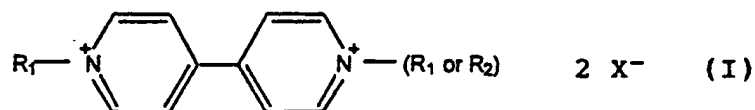
30 Figure 9 : (a) Change in transmittance at 600 nm of
modified EC system 10 containing 0.05, 0.10 and 0.20
 $\text{mol} \cdot \text{dm}^{-3}$ ferrocene during 10 000 test cycles. (b)
Change in colouring and clearing times of modified EC
35 systems 10 in (a) during 10 000 test cycles.

- 17 -

CLAIMS:

1. A nanoporous-nanocrystalline film comprising a
semiconducting metallic oxide having a redox
5 chromophore adsorbed thereto.
2. A film according to claim 1, wherein the redox
chromophore comprises a compound of the general formula
I

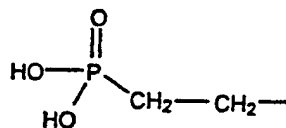
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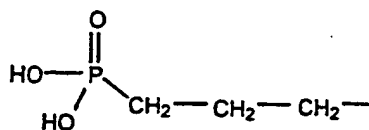
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wherein X is a charge balancing ion such as a
halide; R₁ is any one of the following:

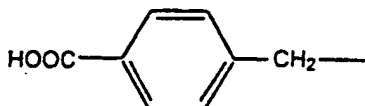
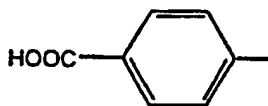
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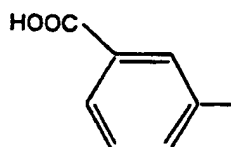
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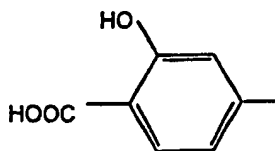
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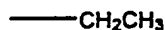
5 and

 R_2 is any one of the following:

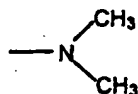
10 (a)



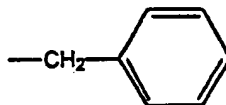
(b)



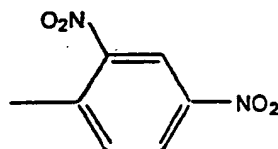
15 (c)



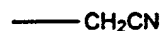
(d)



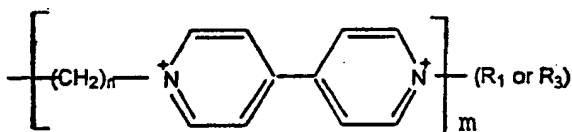
20 (e)



25 (f)



30 (g)

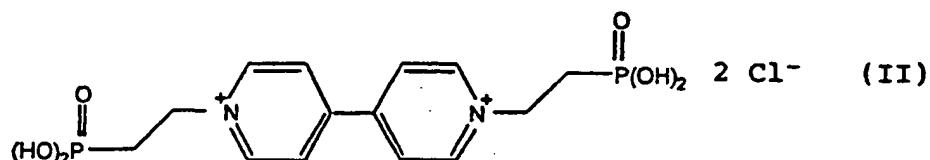


wherein R_1 is as defined above, R_3 is any of the formulae (a) to (f) given above under R_2 , m is an integer of from 1 to 6, preferably 1 or 2 and n is an integer of from 1 to 10.

- 19 -

3. A film according to claim 2, wherein the redox chromophore is a compound of formula II, viz:
bis-(2-phosphonoethyl)-4,4'-bipyridinium dichloride.

5



10 4. A film according to any of claims 1 to 3, wherein the metallic oxide is an oxide of a metal selected from titanium, zirconium, hafnium, chromium, molybdenum, tungsten, vanadium, niobium, tantalum, silver, zinc, strontium, iron (Fe²⁺ and Fe³⁺) and nickel and
15 perovskites thereof, the metallic oxide preferably being selected from TiO₂, WO₃, MoO₃, ZnO and SnO₂.

5. Use of a film according to any of claims 1 to 4 in the manufacture of an electrochromic system.

20

6. An electrochromic system comprising: a first electrode disposed on a transparent or translucent substrate; a second electrode; an electrolyte; an electron donor; and a film according to any of claims 1
25 to 4 intermediate the first and second electrodes.

7. An electrochromic system according to claim 6, wherein the electrolyte is in liquid form, preferably comprising at least one electrochemically inert salt
30 optionally in molten form in solution in a solvent.

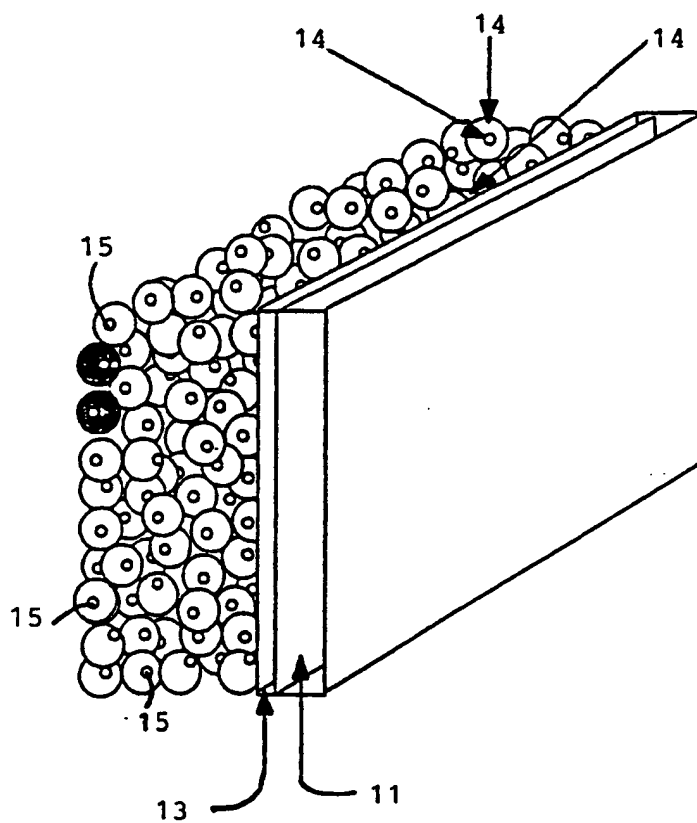
8. An electrochromic system according to claim 7, wherein the salt is selected from hexafluorophosphate, bis-trifluoromethanesulfonate, bis-trifluoromethyl-sulfonylamidure, tetraalkylammonium, dialkyl-1,3-
35

- 20 -

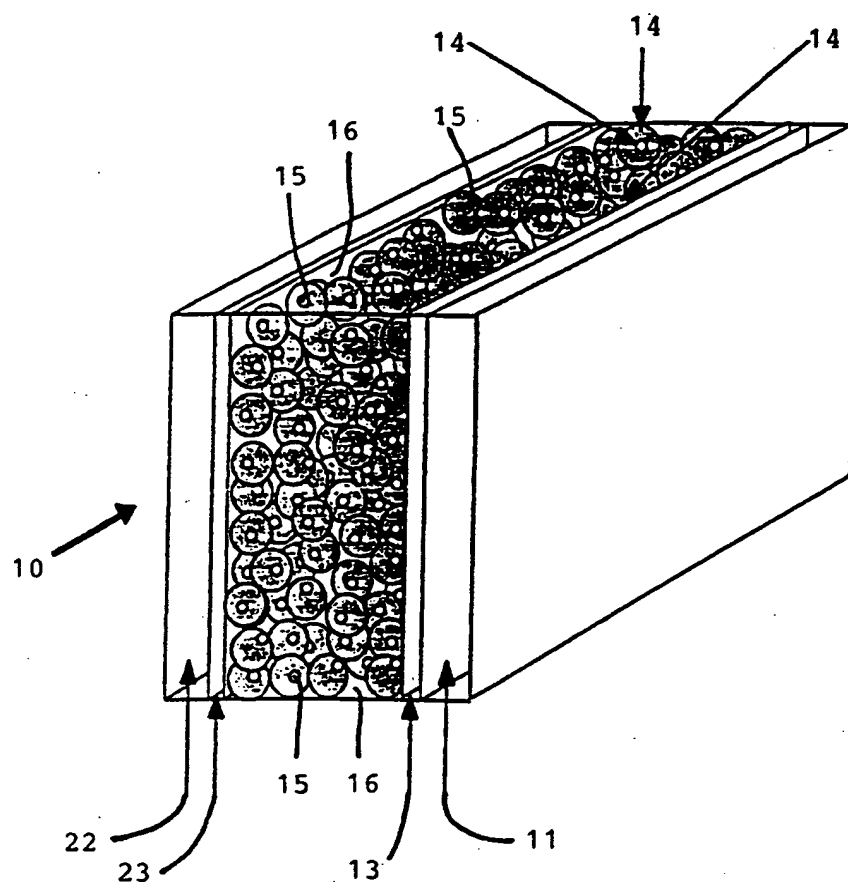
imidazolium and lithium perchlorate or a mixture thereof and/or is a salt in molten form selected from trifluoromethanesulfonate, 1-ethyl,3-methyl imidazolium bis-trifluoromethyl-sulfonylamidure and 1-propyl-dimethyl
5 imidazolium bis-trifluoromethyl-sulfonylamidure or a mixture thereof, the preferred salt being lithium perchlorate; and/or the solvent is selected from acetonitrile, butyronitrile, glutaronitrile, dimethyl-
10 sulfoxide, dimethylformamide, dimethylacetamide, N-methyloxazolidinone, dimethyl-tetrahydropyrimidinone and γ -butyrolactone or a mixture thereof.

9. An electrochromic system according to any of claims
15 6 to 8, wherein the electron donor is a metallocene, preferably ferrocene, or a derivative thereof.

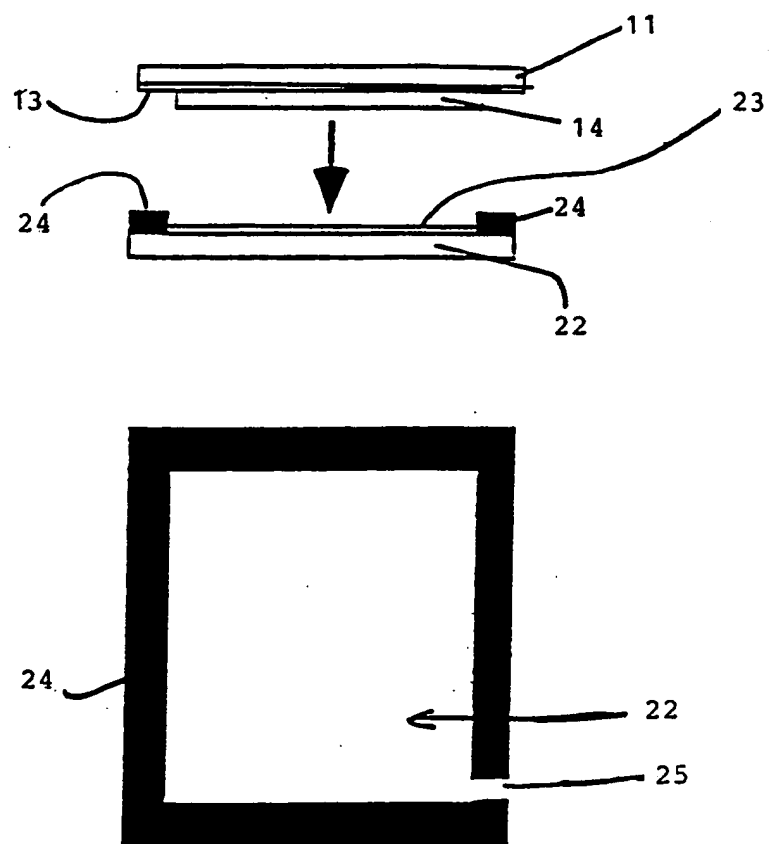
10. An electrochromic system according to claim 9,
20 wherein the ferrocene is in solution in the electrolyte solvent.

1/9**FIG. 1**

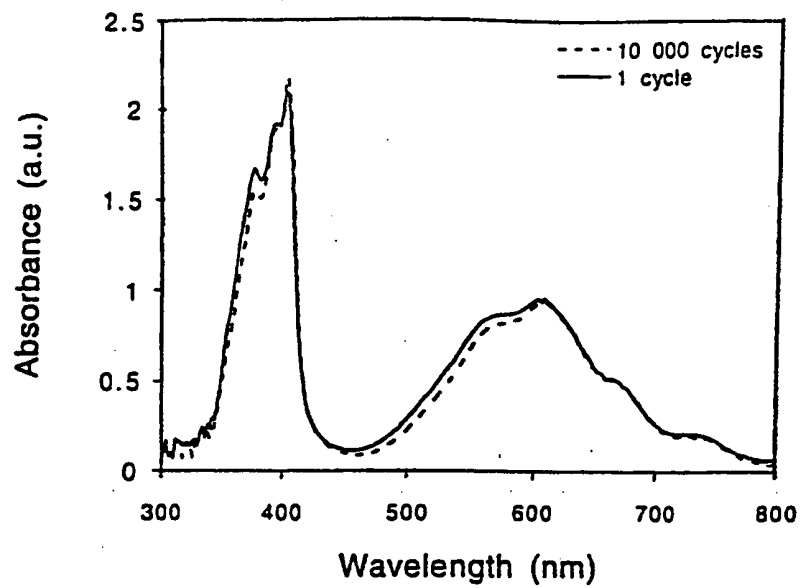
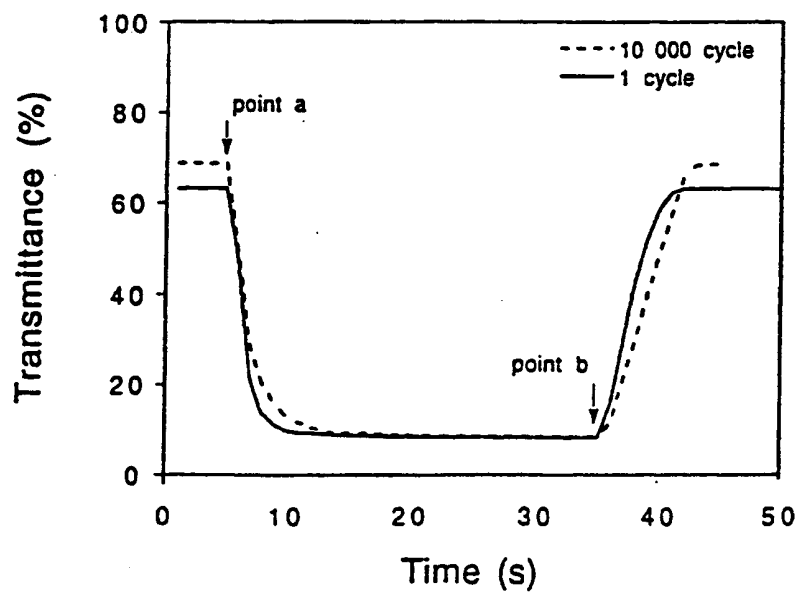
2/9

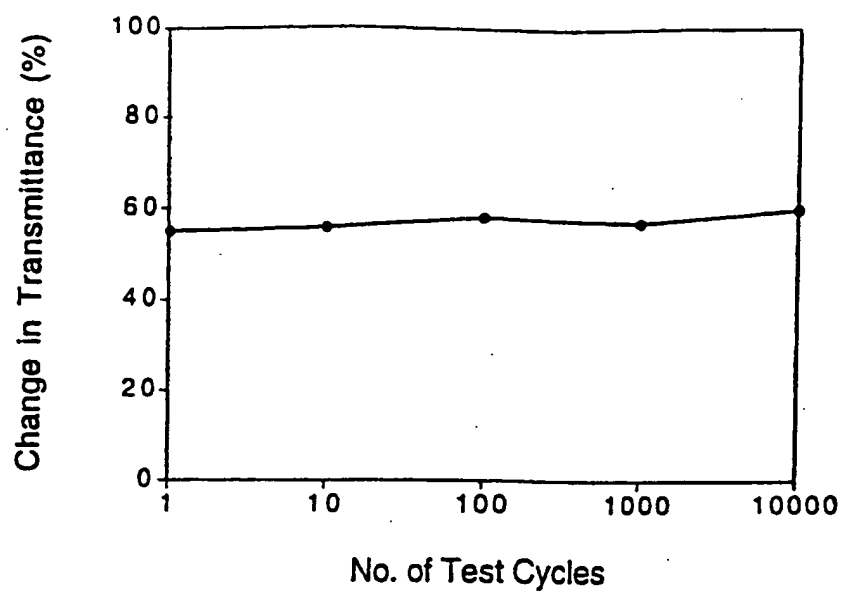
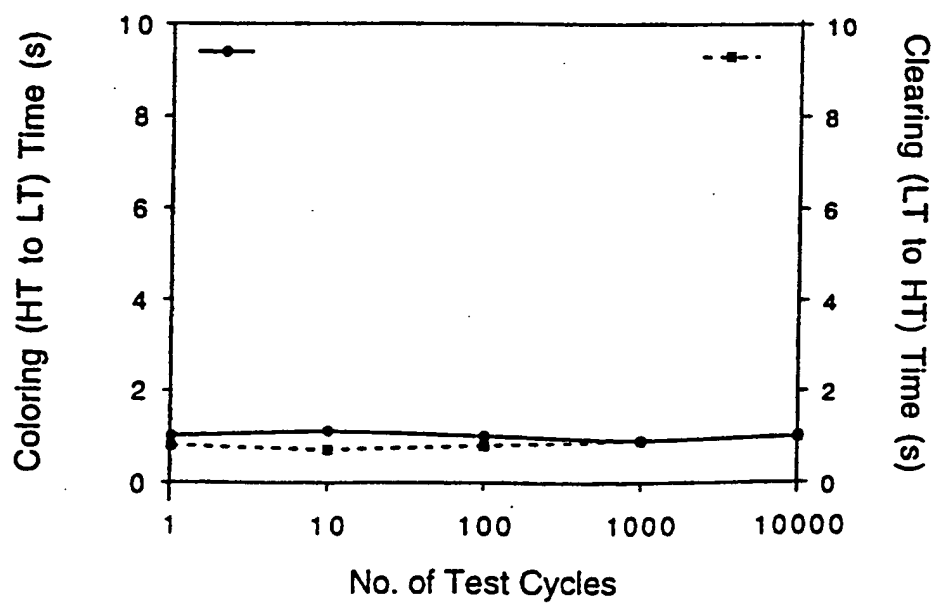
**FIG. 2**

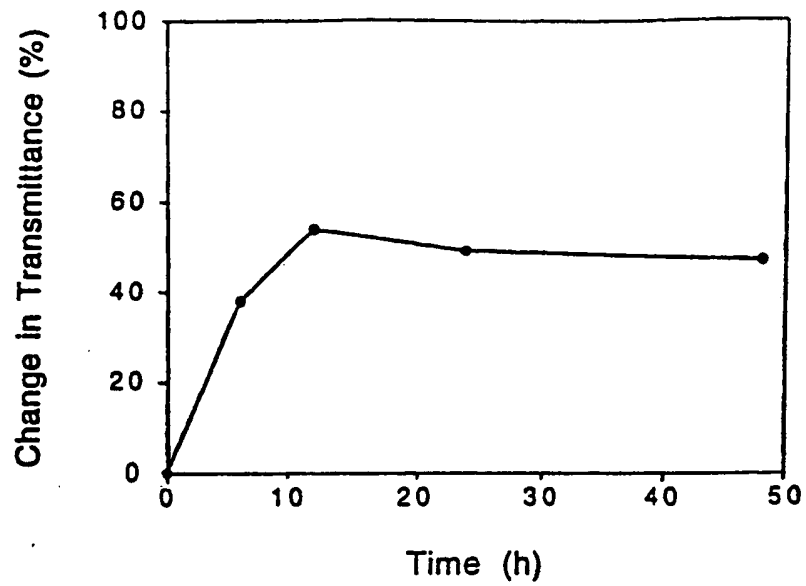
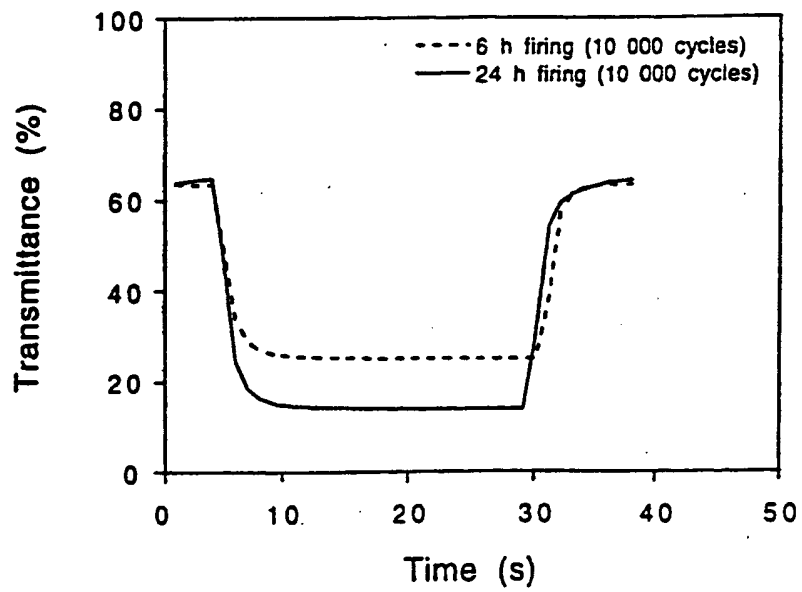
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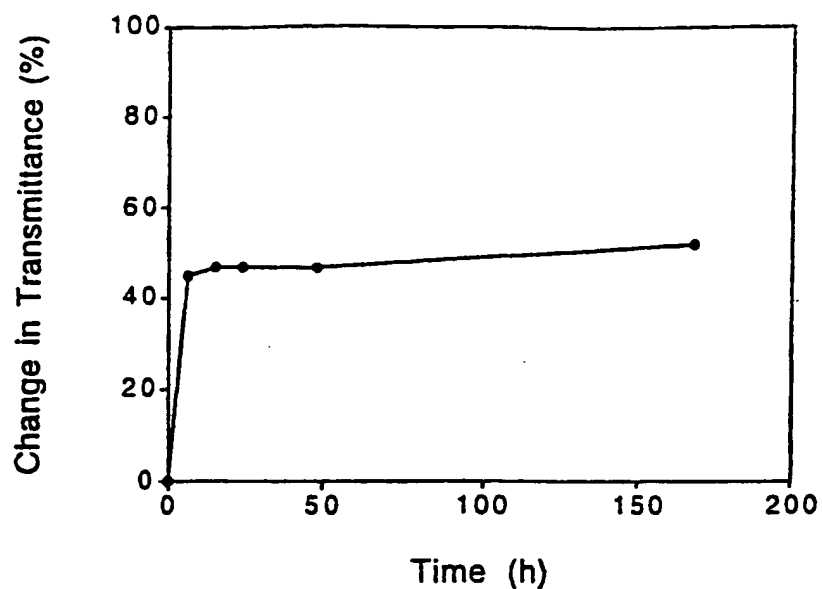
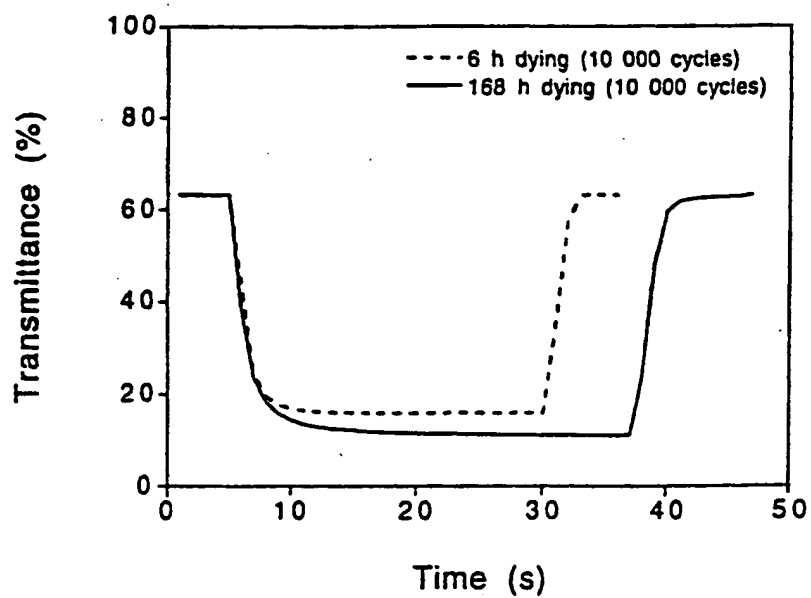
3/9**FIG. 3**

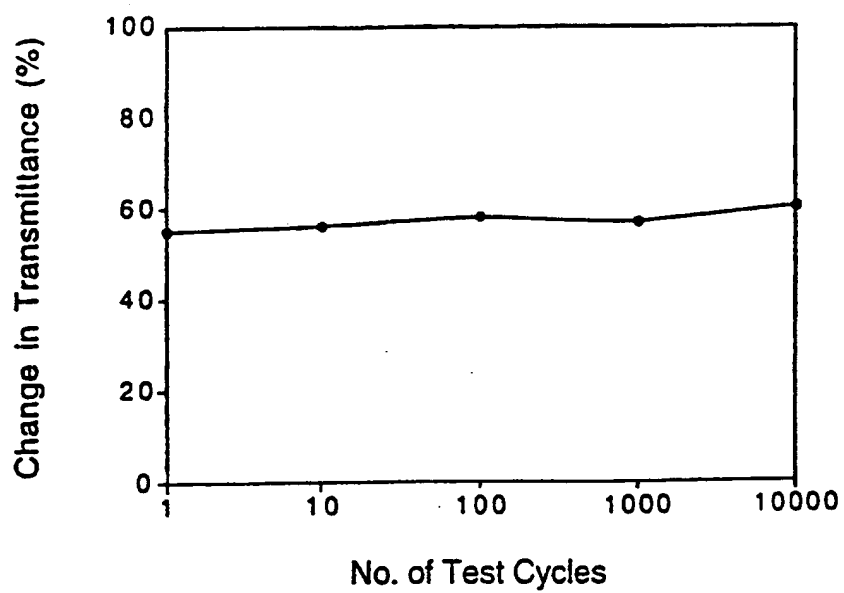
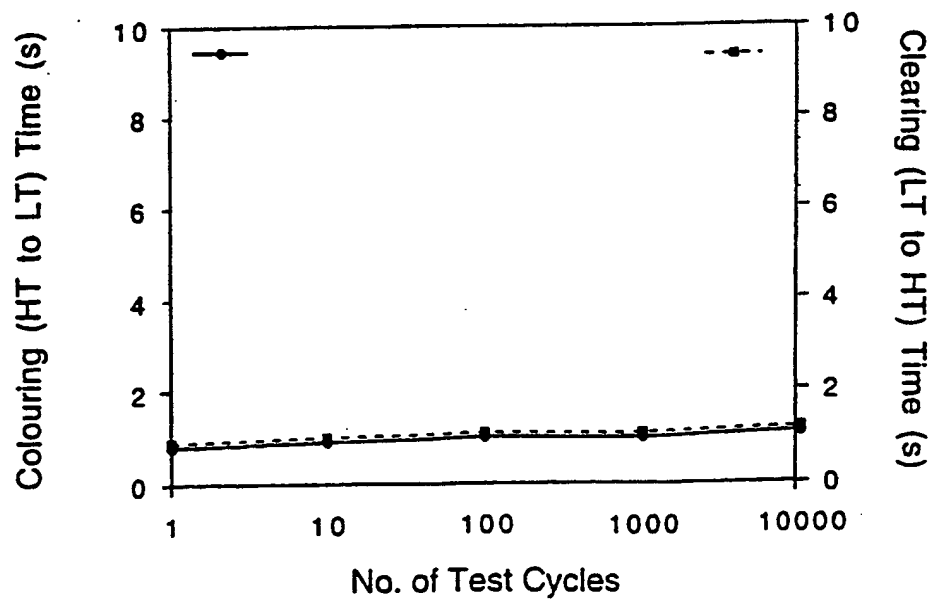
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4/9**FIG. 4a****FIG. 4b**

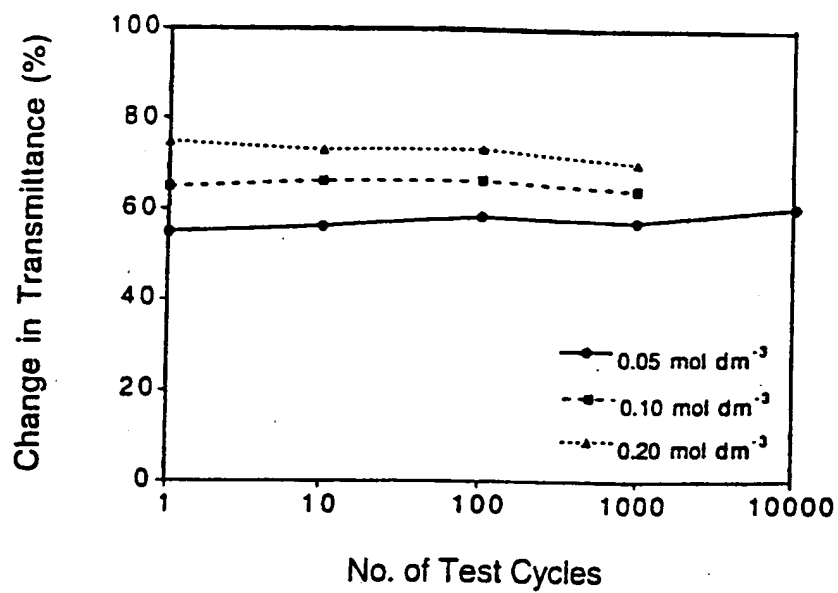
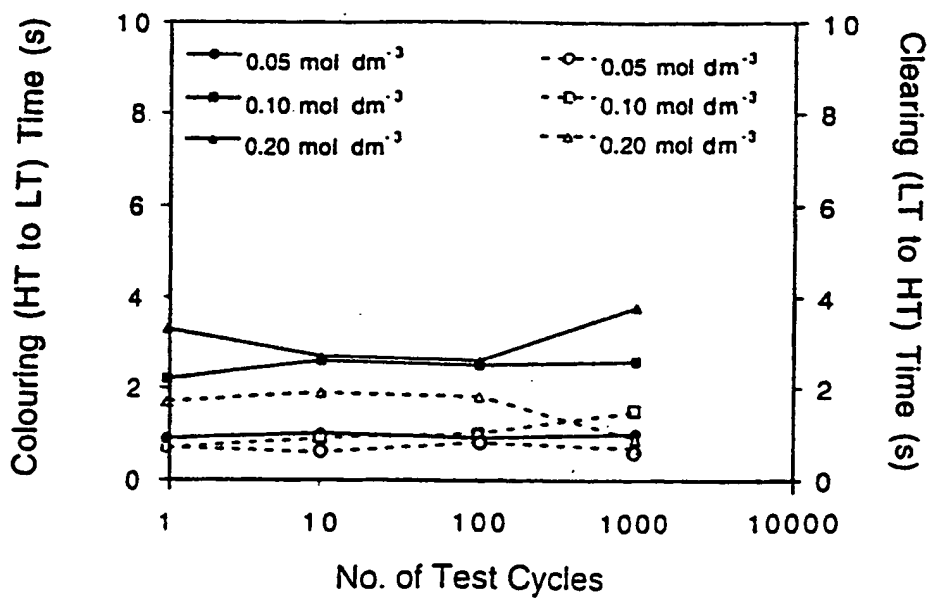
5/9**FIG. 5a****FIG. 5b**

6/9**FIG. 6a****FIG. 6b**

7/9**FIG. 7a****FIG. 7b**

8/9**FIG. 8a****FIG. 8b**

9/9

**FIG. 9a****FIG. 9b**

INTERNATIONAL SEARCH REPORT

International Application No

PCT/IE 98/00008

A. CLASSIFICATION OF SUBJECT MATTER
IPC 6 G02F1/15 C09K9/02

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 G02F C09K

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	BECHINGER C ET AL: "Photoelectrochromic windows and displays" NATURE, 17 OCT. 1996, MACMILLAN MAGAZINES, UK, vol. 383, no. 6601, ISSN 0028-0836, pages 608-610, XP002064353	1,4-6
Y	see the whole document	7-10
P,X	WO 97 35227 A (BONHÔTE) 25 September 1997 see page 13, line 01 - page 15, line 12; figure 10	1,2,4,5
Y	WO 95 30495 A (DONNELLY) 16 November 1995 see page 20, line 08 - page 27, line 02	7-10
	-/--	

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

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- "&" document member of the same patent family

Date of the actual completion of the international search

7 May 1998

Date of mailing of the international search report

02/06/1998

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Diot, P

INTERNATIONAL SEARCH REPORT

International Application No

PCT/IE 98/00008

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	<p>HAGFELD A ET AL: "NANOSTRUCTURED TiO₂ SEMICONDUCTOR ELECTRODES MODIFIED WITH SURFACE ATTACHED VIOLOGENS: APPLICATIONS FOR DISPLAYS AND SMART WINDOWS"</p> <p>PROCEEDINGS OF THE SPIE, vol. 2531, 12 July 1995, pages 60-69, XP000671702 see the whole document</p> <p>-----</p>	1

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/IE 98/00008

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
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WO 9530495 A	16-11-1995	US 5668663 A	16-09-1997
		AU 2463795 A	29-11-1995
		EP 0758929 A	26-02-1997
		JP 10500225 T	06-01-1998
		US 5724187 A	03-03-1998